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Notes

This booklet cannot be used for paper 1 of the examination (HLP1 and SLP1), but the periodic table given on page 4 will be available as part of these examination papers. Clean copies of this booklet must be made available to candidates for papers 2 and 3 (HLP2, HLP3, SLP2 and SLP3).

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1. Some Relevant Equations

$$E = hf$$

$$E = mc^2$$

$$\Delta G^\ominus = \Delta H^\ominus - T\Delta S^\ominus$$

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

$$t_{\frac{1}{2}} = \frac{0.693}{k}$$

$$\log_{10} \frac{I_0}{I} = \epsilon lc$$

$$k = Ae^{\frac{-E_a}{RT}} \quad \ln k = -\frac{E_a}{RT} + \ln A$$

$$[A] = [A]_0 e^{-kt} \quad \text{or} \quad \ln \frac{[A]_0}{[A]} = kt$$

2. Physical Constants

Avogadro's constant (L) = $6.02 \times 10^{23} \text{ mol}^{-1}$

Gas constant (R) = $8.31 \text{ J K}^{-1} \text{ mol}^{-1}$

Molar volume of an ideal gas at 273 K, $1.01 \times 10^5 \text{ Pa} = 2.24 \times 10^{-2} \text{ m}^3 \text{ mol}^{-1}$ ($22.4 \text{ dm}^3 \text{ mol}^{-1}$)

Speed of light in a vacuum (c) = $3.00 \times 10^8 \text{ m s}^{-1}$

Planck constant (h) = $6.63 \times 10^{-34} \text{ J s}$

Specific heat capacity of water = $4.18 \text{ kJ kg}^{-1} \text{ K}^{-1}$ (= $4.18 \text{ J g}^{-1} \text{ K}^{-1}$)

Ionic product constant for water (K_w) = $1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ at 298 K

(1 atm = $1.01 \times 10^5 \text{ Pa}$)

(1 $\text{dm}^3 = 1 \text{ litre} = 1 \times 10^{-3} \text{ m}^3 = 1 \times 10^3 \text{ cm}^3$)

3. Fundamental Particles

	Proton	Neutron	Electron
Mass/kg	1.672648×10^{-27}	1.674954×10^{-27}	9.109534×10^{-31}
Charge/C	1.602189×10^{-19}	0	1.602189×10^{-19}

4. Names of the First 103 Elements

Element	Symbol	Atomic Number	Element	Symbol	Atomic Number
actinium	Ac	89	mercury	Hg	80
aluminium	Al	13	molybdenum	Mo	42
americium	Am	95	neodymium	Nd	60
antimony	Sb	51	neon	Ne	10
argon	Ar	18	neptunium	Np	93
arsenic	As	33	nickel	Ni	28
astatine	At	85	niobium	Nb	41
barium	Ba	56	nitrogen	N	7
berkelium	Bk	97	nobelium	No	102
beryllium	Be	4	osmium	Os	76
bismuth	Bi	83	oxygen	O	8
boron	B	5	palladium	Pd	46
bromine	Br	35	phosphorus	P	15
cadmium	Cd	48	platinum	Pt	78
caesium	Cs	55	plutonium	Pu	94
calcium	Ca	20	polonium	Po	84
californium	Cf	98	potassium	K	19
carbon	C	6	praseodymium	Pr	59
cerium	Ce	58	promethium	Pm	61
chlorine	Cl	17	protactinium	Pa	91
chromium	Cr	24	radium	Ra	88
cobalt	Co	27	radon	Rn	86
copper	Cu	29	rhenium	Re	75
curium	Cm	96	rhodium	Rh	45
dysprosium	Dy	66	rubidium	Rb	37
einsteinium	Es	99	ruthenium	Ru	44
erbium	Er	68	samarium	Sm	62
europium	Eu	63	scandium	Sc	21
fermium	Fm	100	selenium	Se	34
fluorine	F	9	silicon	Si	14
francium	Fr	87	silver	Ag	47
gadolinium	Gd	64	sodium	Na	11
gallium	Ga	31	strontium	Sr	38
germanium	Ge	32	sulfur	S	16
gold	Au	79	tantalum	Ta	73
hafnium	Hf	72	technetium	Tc	43
helium	He	2	tellurium	Te	52
holmium	Ho	67	terbium	Tb	65
hydrogen	H	1	thallium	Tl	81
indium	In	49	thorium	Th	90
iodine	I	53	thulium	Tm	69
iridium	Ir	77	tin	Sn	50
iron	Fe	26	titanium	Ti	22
krypton	Kr	36	tungsten	W	74
lanthanum	La	57	uranium	U	92
lawrencium	Lr	103	vanadium	V	23
lead	Pb	82	xenon	Xe	54
lithium	Li	3	ytterbium	Yb	70
lutetium	Lu	71	yttrium	Y	39
magnesium	Mg	12	zinc	Zn	30
manganese	Mn	25	zirconium	Zr	40
mendelevium	Md	101			

6. Melting Points and Boiling Points of the Elements

		M.pt/K		Element		B.pt/K	
14	H						
20							
454	Li	1551	Be	3243			
1600							
371	Na	922	Mg	1363			
1156							
337	K	1112	Ca	1757	1814	Sc	3104
1047							
312	Rb	1042	Sr	1657	1780	Y	3611
961							
302	Cs	983	Ba	2023	1194	La	3730
952							
300	Fr	973	Ra	1413	1320	Ac	3470
950							
1	He						
4							
25	Ne	54	F	85			
27							
55	O	63	N	77			
392	S	317	P	553			
718							
936	Al	1683	Si	2628			
2740							
303	Ga	1211	Ge	3103			
2676							
693	Zn	63	Cu	1357			
1180							
594	Cd	1726	Ni	3005			
1038							
1235	Ag	1768	Co	3143			
2485							
1825	Pd	1808	Fe	3023			
3413							
2239	Rh	1517	Mn	2235			
4000							
2583	Ru	2130	Cr	2755			
4173							
2445	Tc	1973	V	3650			
5150							
3453	Re	1933	Ti	3560			
5900							
3680	W	1814	Sc	3104			
5930							
3269	Ta	1973	V	3650			
5698							
2503	Hf	1933	Ti	3560			
5470							
1194	La	1780	Y	3611			
3730							
234	Hg	1726	Ni	3005			
630							
1338	Au	1357	Cu	2840			
3080							
2045	Pt	1768	Co	3143			
4100							
2683	Ir	1808	Fe	3023			
4403							
3327	Os	1517	Mn	2235			
5300							
577	Tl	1211	Ge	3103			
1730							
601	Pb	1211	Ge	3103			
2013							
545	Bi	889	As	958			
1833							
527	Po	490	Se	958			
1235							
575	At	266	Br	332			
610							
202	Rn	117	Kr	121			
211							
161	Xe	161	Xe	166			
387	I	904	Sb	2023			
458							
723	Te	723	Te	1263			
1263							

8. Atomic and Ionic Radii of the Elements

		Atomic radius/ 10^{-12} m		Element		Ionic radius/ 10^{-12} m	
		Element	Atomic radius/ 10^{-12} m	Element	Ionic radius/ 10^{-12} m		
30	H	154 (1-)					
152	Li	112					
68 (1+)	Be	30 (2+)					
186	Na	160					
98 (1+)	Mg	65 (2+)					
231	K	197					
133 (1+)	Ca	94 (2+)					
244	Rb	215					
148 (1+)	Sr	110 (2+)					
262	Cs	217					
167 (1+)	Ba	34 (2+)					
270	Fr	220					
88	B	16 (3+)					
143	Al	45 (3+)					
77	C	260 (4-)					
88	N	171 (3-)					
66	O	146 (2-)					
58	F	133 (1-)					
99	Ne	99					
141	Si	42 (4+) 271 (4-)					
122	P	212 (3-)					
117	S	190 (2-)					
114	Cl	181 (1-)					
141	Ar	141					
141	Ga	62 (3+)					
122	Ge	53 (4+) 272 (4-)					
121	As	222 (3-)					
117	Se	202 (2-)					
133	Br	196 (1-)					
166	In	81 (3+)					
162	Sn	112 (2+) 71 (4+)					
141	Sb	245 (3-)					
137	Te	222 (2-)					
133	I	219 (1-)					
175	Xe	175					
171	Pb	120 (2+) 84 (4+)					
170	Bi	120 (3+)					
140	Po	140					
140	At	140					
149	Zn	74 (2+)					
128	Cu	96 (1+) 69 (2+)					
124	Ni	72 (2+)					
138	Ag	126 (1+)					
144	Au	137 (1+) 85 (3+)					
138	Pd						
135	Rh	86 (2+)					
134	Ru	65 (4+)					
138	Pt						
134	Ir	66 (4+)					
135	Os	67 (4+)					
137	Re						
136	Mo	68 (4+)					
137	W	68 (4+)					
141	Nb	70 (5+)					
143	Ta	73 (5+)					
157	Zr	80 (4+)					
157	Hf	81 (4+)					
180	Y	93 (3+)					
188	La	115 (3+)					
200	Ac	200					
160	Sc	81 (3+)					
146	Ti	90 (2+) 68 (4+)					
146	V	88 (2+) 59 (5+)					
131	Cr	63 (3+)					
125	Mn	80 (2+) 60 (4+)					
126	Fe	76 (2+) 64 (3+)					
125	Co	74 (2+) 63 (3+)					
124	Ni	72 (2+)					
133	Zn	74 (2+)					

9. Covalent Bond Lengths

Bond	Bond length /nm	Bond	Bond length /nm
H-H	0.074	C-H	0.109
C-C	0.154	Si-H	0.146
C=C	0.134	N-H	0.101
C≡C	0.120	P-H	0.142
C-C (in benzene)	0.139	O-H	0.096
Si-Si	0.235	S-H	0.135
N-N	0.146	F-H	0.092
N=N	0.120	Cl-H	0.128
N≡N	0.110	Br-H	0.141
P-P (P ₄)	0.221	I-H	0.160
O-O	0.148		
O=O	0.121	C-O	0.143
S-S (S ₈)	0.207	C=O	0.122
S=S	0.188	C-O (in phenol)	0.136
F-F	0.142	C-N	0.147
Cl-Cl	0.199	C=N	0.127
Br-Br	0.228	C≡N	0.116
I-I	0.267	C-N (in phenylamine)	0.135
		C-F	0.138
		C-Cl	0.177
		C-Cl (in chlorobenzene)	0.169
		C-Br	0.193
		C-I	0.214
		Si-O	0.150

10. Average Bond Enthalpies at 298 K

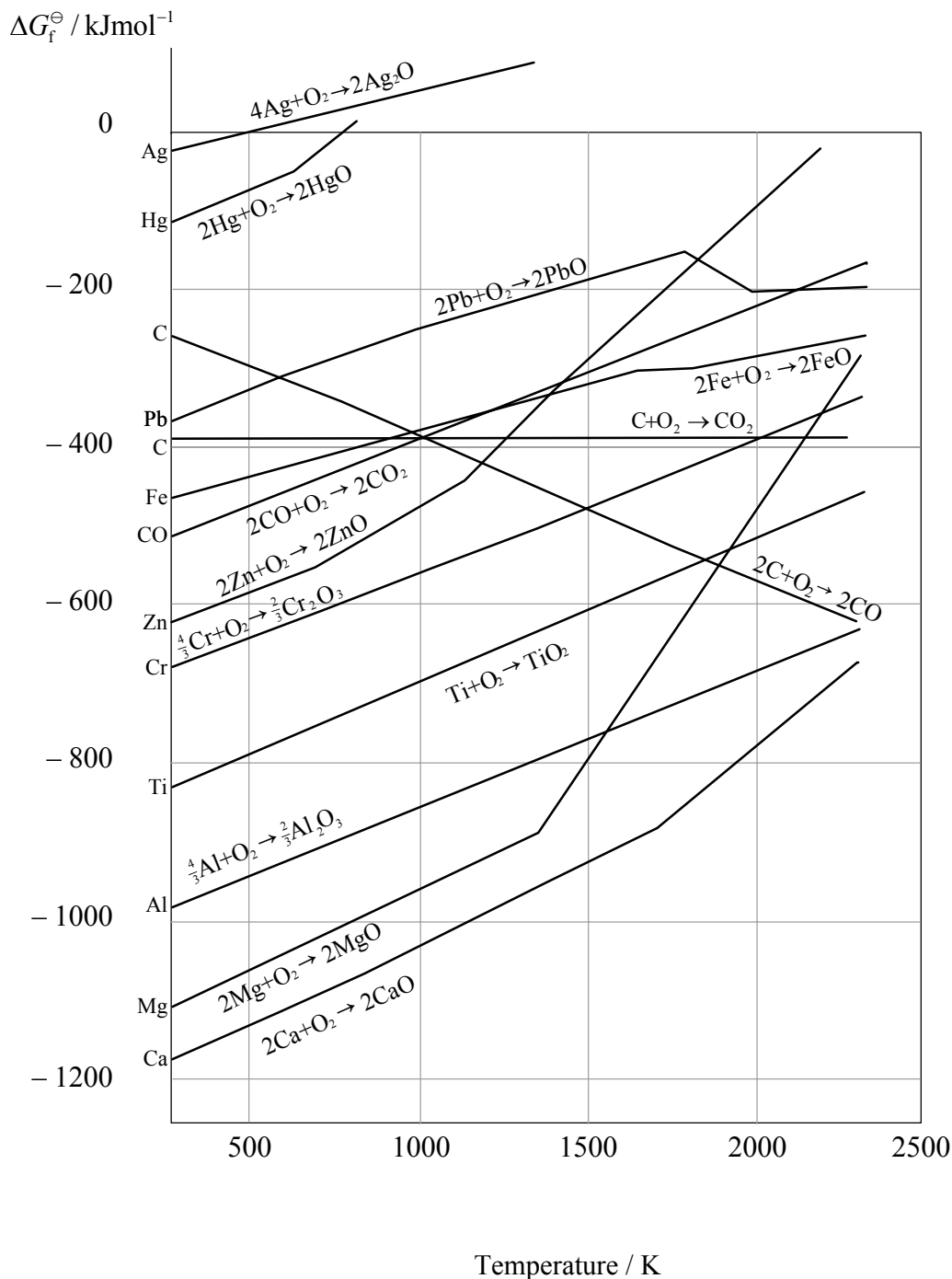
Bond	$\Delta H / \text{kJ mol}^{-1}$	Bond	$\Delta H / \text{kJ mol}^{-1}$
H-H	436	C-H	412
D-D	442	Si-H	318
C-C	348	N-H	388
C=C	612	P-H	322
C≡C	837	O-H	463
C-C (benzene)	518	S-H	338
Si-Si	226	F-H	562
Ge-Ge	188	Cl-H	431
Sn-Sn	151	Br-H	366
N-N	163	I-H	299
N=N	409		
N≡N	944	C-O	360
P-P	172	C=O	743
O-O	146	C-N	305
O=O	496	C=N	613
S-S	264	C≡N	890
F-F	158	C-F	484
Cl-Cl	242	C-Cl	338
Br-Br	193	C-Br	276
I-I	151	C-I	238
		Si-O	374

11. Organic Compounds—Thermodynamic Data

Substance	Formula	State	$\Delta H_f^\ominus / \text{kJ mol}^{-1}$	$\Delta G_f^\ominus / \text{kJ mol}^{-1}$	$S^\ominus / \text{J K}^{-1} \text{mol}^{-1}$
methane	CH ₄	g	-75	-51	186
ethane	C ₂ H ₆	g	-85	-33	230
propane	C ₃ H ₈	g	-104	-24	270
butane	C ₄ H ₁₀	g	-125	-16	310
pentane	C ₅ H ₁₂	g	-146	-8	348
hexane	C ₆ H ₁₄	g	-167	0	387
ethene	C ₂ H ₄	g	52	68	219
propene	C ₃ H ₆	g	20	63	267
but-1-ene	C ₄ H ₈	g	1	72	307
<i>cis</i> -but-2-ene	C ₄ H ₈	g	-6	67	301
<i>trans</i> -but-2-ene	C ₄ H ₈	g	-10	64	296
ethyne	C ₂ H ₂	g	227	209	201
propyne	C ₃ H ₄	g	185	194	248
buta-1,3-diene	C ₄ H ₆	g	112	152	279
cyclohexane	C ₆ H ₁₂	l	-156	27	204
benzene	C ₆ H ₆	g	83	130	269
benzene	C ₆ H ₆	l	49	125	173
methylbenzene	C ₆ H ₅ CH ₃	g	50	122	320
ethylbenzene	C ₆ H ₅ CH ₂ CH ₃	g	30	131	360
phenylethene	C ₆ H ₅ CHCH ₂	g	148	214	345
chloromethane	CH ₃ Cl	g	-82	-59	234
dichloromethane	CH ₂ Cl ₂	l	-117	-63	179
trichloromethane	CHCl ₃	l	-132	-72	203
bromomethane	CH ₃ Br	g	-36	-26	246
tribromomethane	CHBr ₃	l	-20	3	222
iodomethane	CH ₃ I	l	-8	20	163
triiodomethane	CHI ₃	s	141		
chloroethane	C ₂ H ₅ Cl	g	-105	-53	276
bromoethane	C ₂ H ₅ Br	l	-85		
iodoethane	C ₂ H ₅ I	l	-31		
chloroethene	C ₂ H ₃ Cl	g	31	52	264
1,2-dichloroethane	CH ₂ ClCH ₂ Cl	l	-166	-80	208
chlorobenzene	C ₆ H ₅ Cl	g	52	99	314
methanol	CH ₃ OH	g	-201	-162	238
methanol	CH ₃ OH	l	-239	-166	127
ethanol	C ₂ H ₅ OH	g	-235	-169	282
ethanol	C ₂ H ₅ OH	l	-278	-175	161
phenol	C ₆ H ₅ OH	s	-163	-51	146
methanal	HCHO	g	-116	-110	219
ethanal	CH ₃ CHO	g	-166	-134	266
propanone	(CH ₃) ₂ CO	g	-216	-152	295
methanoic acid	HCOOH	l	-409	-346	129
ethanoic acid	CH ₃ COOH	l	-487	-392	160
benzoic acid	C ₆ H ₅ COOH	s	-385	-245	167
ethyl ethanoate	CH ₃ COOC ₂ H ₅	l	-481		
ethanamide	CH ₃ CONH ₂	s	-320		
methylamine	CH ₃ NH ₂	g	-28	28	242
ethylamine	C ₂ H ₅ NH ₂	g	-49	37	285
urea	CO(NH ₂) ₂	s	-333	-47	105

12. Ellingham Diagram

Standard Gibbs free energy changes of formation, ΔG_f^\ominus , for oxides given per mole of oxygen gas as a function of temperature.



13. Enthalpies of Combustion

The values of the molar enthalpy of combustion (ΔH_c^\ominus) in the following table refer to a temperature of 298 K and a pressure of 1.01×10^5 Pa (1 atm).

Substance	Formula	State	$\Delta H_c^\ominus / \text{kJ mol}^{-1}$	Substance	Formula	State	$\Delta H_c^\ominus / \text{kJ mol}^{-1}$
hydrogen	H ₂	g	-286	propan-1-ol	C ₃ H ₇ OH	l	-2010
sulfur	S	s	-297	butan-1-ol	C ₄ H ₉ OH	l	-2673
carbon (graphite)	C	s	-394	phenylmethanol	C ₆ H ₅ CH ₂ OH	l	-4056
carbon (diamond)	C	s	-395	cyclohexanol	C ₆ H ₁₁ OH	s	-3727
carbon monoxide	CO	g	-283	phenol	C ₆ H ₅ OH	s	-3064
methane	CH ₄	g	-890	ethoxyethane	(C ₂ H ₅) ₂ O	l	-2727
ethane	C ₂ H ₆	g	-1560	methanal	HCHO	g	-561
propane	C ₃ H ₈	g	-2220	ethanal	CH ₃ CHO	l	-1167
butane	C ₄ H ₁₀	g	-2877	benzaldehyde	C ₆ H ₅ CHO	l	-3520
pentane	C ₅ H ₁₂	g	-3509	propanone	(CH ₃) ₂ CO	l	-1786
hexane	C ₆ H ₁₄	l	-4194	pentan-3-one	(C ₂ H ₅) ₂ CO	l	-3078
octane	C ₈ H ₁₈	l	-5512	phenylethanol	CH ₃ COC ₆ H ₅	s	-4138
cyclohexane	C ₆ H ₁₂	l	-3924	diphenylmethanone	(C ₆ H ₅) ₂ CO	s	-6512
ethene	C ₂ H ₄	l	-1409	methanoic acid	HCOOH	l	-263
buta-1,3-diene	C ₄ H ₆	g	-2542	ethanoic acid	CH ₃ COOH	l	-876
ethyne	C ₂ H ₂	g	-1299	benzoic acid	C ₆ H ₅ COOH	s	-3227
benzene	C ₆ H ₆	g	-3273	ethanedioic acid	(COOH) ₂	s	-246
methylbenzene	C ₆ H ₅ CH ₃	l	-3909	ethyl ethanoate	CH ₃ COOC ₂ H ₅	l	-2246
naphthalene	C ₁₀ H ₈	l	-5157	ethanamide	CH ₃ CONH ₂	s	-1182
anthracene	C ₁₄ H ₁₀	s	-7114	benzamide	C ₆ H ₅ CONH ₂	s	-3546
chloroethane	C ₂ H ₅ Cl	s	-1325	methylamine	CH ₃ NH ₂	g	-1072
bromoethane	C ₂ H ₅ Br	g	-1425	ethylamine	C ₂ H ₅ NH ₂	g	-1709
iodoethane	C ₂ H ₅ I	g	-1490	phenylamine	C ₆ H ₅ NH ₂	l	-3397
(chloromethyl)benzene	C ₆ H ₅ CH ₂ Cl	l	-3709	nitrobenzene	C ₆ H ₅ NO ₂	l	-3094
trichloromethane	CHCl ₃	l	-373	urea	CO(NH ₂) ₂	s	-634
methanol	CH ₃ OH	l	-715	glucose	C ₆ H ₁₂ O ₆	s	-2816
ethanol	C ₂ H ₅ OH	l	-1371	sucrose	C ₁₂ H ₂₂ O ₁₁	s	-5644

14. Lattice Enthalpies at 298 K (Experimental and Theoretical Values)

The lattice enthalpy ($\Delta H_{\text{lattice}}^{\ominus}$) values given relate to the endothermic process $\text{MX}(\text{s}) \rightarrow \text{M}^+(\text{g}) + \text{X}^-(\text{g})$ in which the gaseous ions of a crystal are separated to an infinite distance from each other.

Experimental Values

The data in these two tables are experimental values obtained by means of a suitable Born–Haber cycle.

Alkali metal halides	$\Delta H_{\text{lattice}}^{\ominus} / \text{kJ mol}^{-1}$			
	F	Cl	Br	I
Li	1022	846	800	744
Na	902	771	733	684
K	801	701	670	629
Rb	767	675	647	609
Cs	716	645	619	585

Other substances	$\Delta H_{\text{lattice}}^{\ominus} / \text{kJ mol}^{-1}$	Other substances	$\Delta H_{\text{lattice}}^{\ominus} / \text{kJ mol}^{-1}$
CaF ₂	2602	MgS	3238
BeCl ₂	3006	CaS	2966
MgCl ₂	2493	SrS	2779
CaCl ₂	2237	BaS	2643
SrCl ₂	2112	CuCl	976
BaCl ₂	2018	AgF	955
MgO	3889	AgCl	905
CaO	3513	AgBr	890
SrO	3310	AgI	876
BaO	3152	NH ₄ Cl	640

Theoretical Values

These two tables contain lattice enthalpies calculated from electrostatic principles on the basis of a purely ionic model for the crystal.

Alkali metal halides	$\Delta H_{\text{lattice}}^{\ominus} / \text{kJ mol}^{-1}$			
	F	Cl	Br	I
Li	1004	833	787	728
Na	891	766	732	686
K	795	690	665	632
Rb	761	674	644	607
Cs	728	636	611	582

Other substances	$\Delta H_{\text{lattice}}^{\ominus} / \text{kJ mol}^{-1}$	Other substances	$\Delta H_{\text{lattice}}^{\ominus} / \text{kJ mol}^{-1}$
CaF ₂	2611	AgF	870
MgO	3929	AgCl	770
CaO	3477	AgBr	758
SrO	3205	AgI	736
BaO	3042		

15. Standard Electrode Potentials

Oxidized species	\rightleftharpoons	Reduced species	E^\ominus / V
$\text{Li}^+(\text{aq}) + \text{e}^-$	\rightleftharpoons	$\text{Li}(\text{s})$	-3.03
$\text{K}^+(\text{aq}) + \text{e}^-$	\rightleftharpoons	$\text{K}(\text{s})$	-2.92
$\text{Ca}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Ca}(\text{s})$	-2.87
$\text{Na}^+(\text{aq}) + \text{e}^-$	\rightleftharpoons	$\text{Na}(\text{s})$	-2.71
$\text{Mg}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Mg}(\text{s})$	-2.36
$\text{Al}^{3+}(\text{aq}) + 3\text{e}^-$	\rightleftharpoons	$\text{Al}(\text{s})$	-1.66
$\text{Mn}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Mn}(\text{s})$	-1.18
$\text{H}_2\text{O}(\text{l}) + \text{e}^-$	\rightleftharpoons	$\frac{1}{2}\text{H}_2(\text{g}) + \text{OH}^-(\text{aq})$	-0.83
$\text{Zn}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Zn}(\text{s})$	-0.76
$\text{Fe}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Fe}(\text{s})$	-0.44
$\text{Ni}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Ni}(\text{s})$	-0.23
$\text{Sn}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Sn}(\text{s})$	-0.14
$\text{Pb}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Pb}(\text{s})$	-0.13
$\text{H}^+(\text{aq}) + \text{e}^-$	\rightleftharpoons	$\frac{1}{2}\text{H}_2(\text{g})$	0.00
$\text{Cu}^{2+}(\text{aq}) + \text{e}^-$	\rightleftharpoons	$\text{Cu}^+(\text{s})$	+0.15
$\text{SO}_4^{2-}(\text{aq}) + 4\text{H}^+(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$	+0.17
$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{Cu}(\text{s})$	+0.34
$\frac{1}{2}\text{O}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) + 2\text{e}^-$	\rightleftharpoons	$2\text{OH}^-(\text{aq})$	+0.40
$\text{Cu}^+(\text{aq}) + \text{e}^-$	\rightleftharpoons	$\text{Cu}(\text{s})$	+0.52
$\frac{1}{2}\text{I}_2(\text{s}) + \text{e}^-$	\rightleftharpoons	$\text{I}^-(\text{aq})$	+0.54
$\text{Fe}^{3+}(\text{aq}) + \text{e}^-$	\rightleftharpoons	$\text{Fe}^{2+}(\text{aq})$	+0.77
$\text{Ag}^+(\text{aq}) + \text{e}^-$	\rightleftharpoons	$\text{Ag}(\text{s})$	+0.80
$\frac{1}{2}\text{Br}_2(\text{l}) + \text{e}^-$	\rightleftharpoons	$\text{Br}^-(\text{aq})$	+1.09
$\frac{1}{2}\text{O}_2(\text{g}) + 2\text{H}^+(\text{aq}) + 2\text{e}^-$	\rightleftharpoons	$\text{H}_2\text{O}(\text{l})$	+1.23
$\text{Cr}_2\text{O}_7^{2-}(\text{aq}) + 14\text{H}^+(\text{aq}) + 6\text{e}^-$	\rightleftharpoons	$2\text{Cr}^{3+}(\text{aq}) + 7\text{H}_2\text{O}(\text{l})$	+1.33
$\frac{1}{2}\text{Cl}_2(\text{g}) + \text{e}^-$	\rightleftharpoons	$\text{Cl}^-(\text{aq})$	+1.36
$\text{MnO}_4^-(\text{aq}) + 8\text{H}^+(\text{aq}) + 5\text{e}^-$	\rightleftharpoons	$\text{Mn}^{2+}(\text{aq}) + 4\text{H}_2\text{O}(\text{l})$	+1.51
$\frac{1}{2}\text{F}_2(\text{g}) + \text{e}^-$	\rightleftharpoons	$\text{F}^-(\text{aq})$	+2.87

16. Strengths of Organic Acids and Bases

The acid strengths in the following tables are given in terms of pK_a values, where $pK_a = -\log_{10} K_a$. The dissociation constant, K_a , values are for aqueous solutions at 298 K. Base strengths are given in terms of pK_b values.

Carboxylic Acids

Name	Formula	pK_a
methanoic	HCOOH	3.75
ethanoic	CH ₃ COOH	4.76
propanoic	CH ₃ CH ₂ COOH	4.87
butanoic	CH ₃ (CH ₂) ₂ COOH	4.82
2-methylpropanoic	(CH ₃) ₂ CHCOOH	4.85
pentanoic	CH ₃ (CH ₂) ₃ COOH	4.86
2,2-dimethylpropanoic	(CH ₃) ₃ CCOOH	5.05
benzoic	C ₆ H ₅ COOH	4.20
phenylethanoic	C ₆ H ₅ CH ₂ COOH	4.31

Halogenated Carboxylic Acids

Name	Formula	pK_a
chloroethanoic	CH ₂ ClCOOH	2.86
dichloroethanoic	CHCl ₂ COOH	1.29
trichloroethanoic	CCl ₃ COOH	0.65
fluoroethanoic	CH ₂ FCOOH	2.66
bromoethanoic	CH ₂ BrCOOH	2.90
iodoethanoic	CH ₂ ICOOH	3.17

Phenols

Name	Formula	pK_a
phenol	C ₆ H ₅ OH	10.00
2-nitrophenol	O ₂ NC ₆ H ₄ OH	7.21
3-nitrophenol	O ₂ NC ₆ H ₄ OH	8.35
4-nitrophenol	O ₂ NC ₆ H ₄ OH	7.15
2,4-dinitrophenol	(O ₂ N) ₂ C ₆ H ₃ OH	4.01
2,4,6-trinitrophenol	(O ₂ N) ₃ C ₆ H ₂ OH	0.42

Alcohols

Name	Formula	pK_a
methanol	CH ₃ OH	15.5
ethanol	C ₂ H ₅ OH	16 (approximately)

Amines

Name	Formula	p <i>K</i> _b
ammonia	NH ₃	4.75
methylamine	CH ₃ NH ₂	3.36
ethylamine	CH ₃ CH ₂ NH ₂	3.27
dimethylamine	(CH ₃) ₂ NH	3.28
trimethylamine	(CH ₃) ₃ N	4.20
diethylamine	(C ₂ H ₅) ₂ NH	3.07
triethylamine	(C ₂ H ₅) ₃ N	3.36
phenylamine	C ₆ H ₅ NH ₂	9.38

17. Acid–base Indicators

			Colour change	
Indicator	p <i>K</i> _a	pH range	Acid	Alkali
methyl orange	3.7	3.1–4.4	red	yellow
bromophenol blue	4.0	3.0–4.6	yellow	blue
bromocresol green	4.7	3.8–5.4	yellow	blue
methyl red	5.1	4.2–6.3	red	yellow
bromothymol blue	7.0	6.0–7.6	yellow	blue
phenol red	7.9	6.8–8.4	yellow	red
phenolphthalein	9.3	8.3–10.0	colourless	red

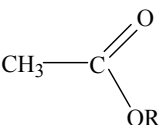
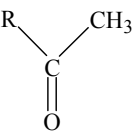
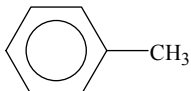
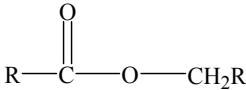
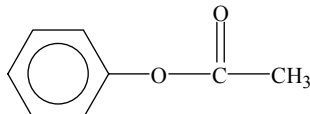
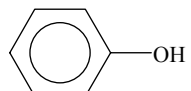
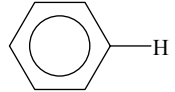
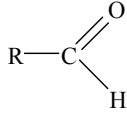
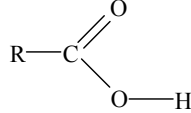
18. Infrared Data

Characteristic ranges for infrared absorption due to stretching vibrations in organic molecules.

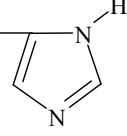
Bond	Organic molecules	Wavenumber / cm⁻¹
C–Cl	halogenoalkanes	700 to 800
C–O	alcohols, ethers, esters	1000 to 1300
C=C	alkenes	1610 to 1680
C=O	aldehydes, ketones, acids, esters	1680 to 1750
C≡C	alkynes	2070 to 2250
O–H	“hydrogen bonded” in acids	2500 to 3300
C–H	alkanes, alkenes, arenes	2840 to 3095
O–H	“hydrogen bonded” in alcohols, phenols	3230 to 3550
N–H	primary amines	3350 to 3500

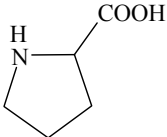
19. ¹H NMR Data

Typical proton chemical shift values (δ) relative to TMS = 0.
(These values can vary slightly in different solvents.)

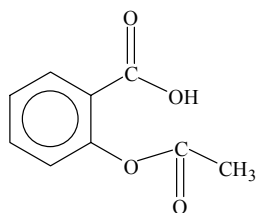
Type of proton	Chemical shift/ppm
R—CH ₃	0.9
R—CH ₂ —R	1.3
R ₃ CH	2
	2.0
	2.1
	2.3
R—C≡C—H	2.6
R—CH ₂ —Hal	3.2–3.7
R—O—CH ₃	3.8
	4.1
	4.0–4.2
R—O—H	0.5–6.5 (can vary considerably under different conditions)
RHC=CH ₂	4.9–5.9
	7
	7.3
	9.7
	11.5

20. 2-amino Acids

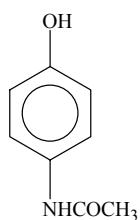
Common name	Symbol	Structural formula	pH of isoelectric point
alanine	Ala	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_3 \end{array}$	6.0
arginine	Arg	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{CH}_2-\text{CH}_2-\text{NH}-\text{C}-\text{NH}_2 \\ \quad \quad \quad \quad \quad \quad \\ \quad \quad \quad \quad \quad \quad \text{NH} \end{array}$	10.8
asparagine	Asn	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{C}-\text{NH}_2 \\ \quad \quad \\ \quad \quad \text{O} \end{array}$	5.4
aspartic acid	Asp	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{COOH} \end{array}$	2.8
cysteine	Cys	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{SH} \end{array}$	5.1
glutamine	Gln	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{CH}_2-\text{C}-\text{NH}_2 \\ \quad \quad \quad \quad \quad \\ \quad \quad \quad \quad \quad \quad \text{O} \end{array}$	5.7
glutamic acid	Glu	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{CH}_2-\text{COOH} \end{array}$	3.2
glycine	Gly	$\text{H}_2\text{N}-\text{CH}_2-\text{COOH}$	6.0
histidine	His	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{C}_4\text{H}_3\text{N}_2 \end{array}$ <p style="text-align: center;">  </p>	7.6
isoleucine	Ile	$\begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{CH}_3 \end{array}$	6.0

Common name	Symbol	Structural formula	pH of isoelectric point
leucine	Leu	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2 \\ \\ \text{CH}_3-\text{CH}-\text{CH}_3 \end{array} $	6.0
lysine	Lys	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{NH}_2 \end{array} $	9.7
methionine	Met	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{CH}_2-\text{S}-\text{CH}_3 \end{array} $	5.7
phenylalanine	Phe	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{C}_6\text{H}_5 \end{array} $	5.5
proline	Pro		6.3
serine	Ser	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{OH} \end{array} $	5.7
threonine	Thr	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_3-\text{CH}-\text{OH} \end{array} $	5.6
tryptophan	Trp	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{C}_8\text{H}_6\text{N}_2 \end{array} $	5.9
tyrosine	Tyr	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_2-\text{C}_6\text{H}_4-\text{OH} \end{array} $	5.7
valine	Val	$ \begin{array}{c} \text{H}_2\text{N}-\text{CH}-\text{COOH} \\ \\ \text{CH}_3-\text{CH}-\text{CH}_3 \end{array} $	6.0

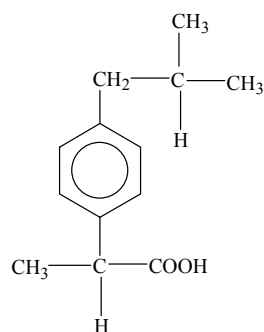
21. Structural Formulas of Some Important Medicines and Drugs



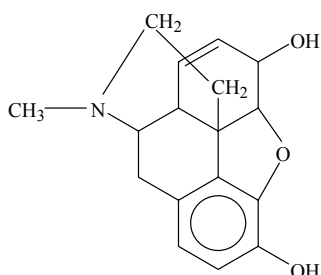
aspirin



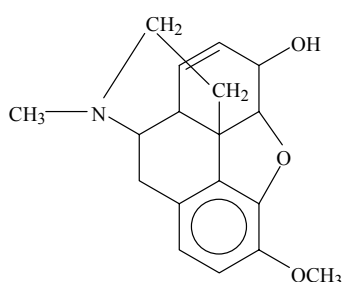
paracetamol (acetaminophen)



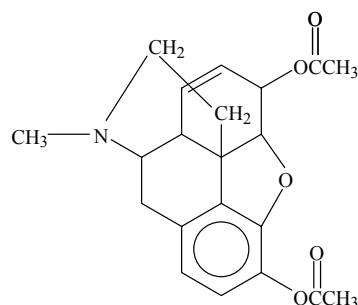
ibuprofen



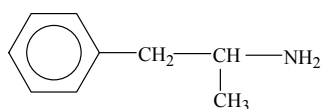
morphine



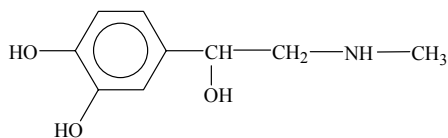
codeine



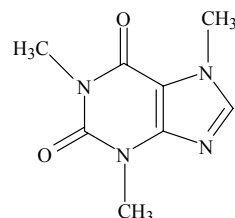
heroin



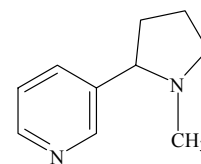
amphetamine



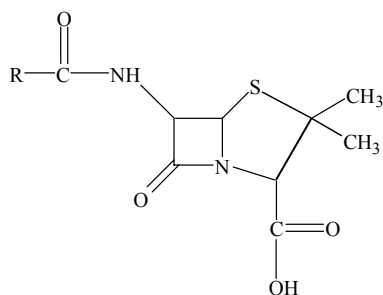
adrenaline



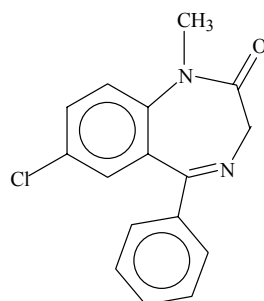
caffeine



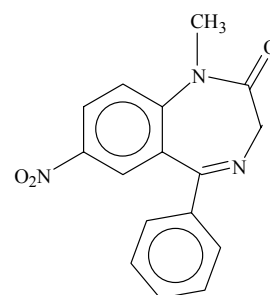
nicotine



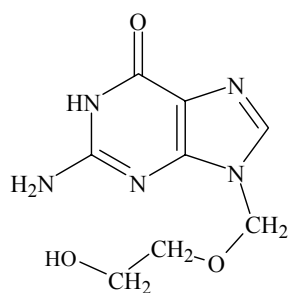
penicillin



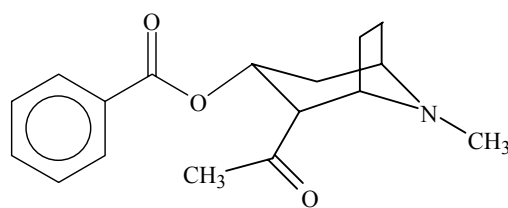
diazepam (valium®)



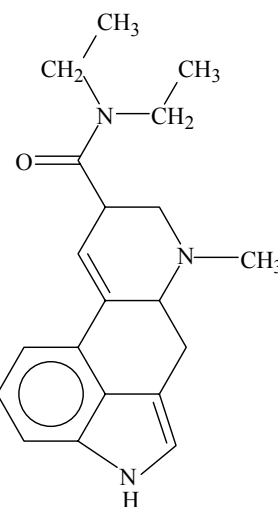
nitrazepam (mogadon®)



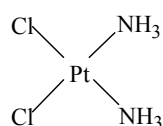
acyclovir



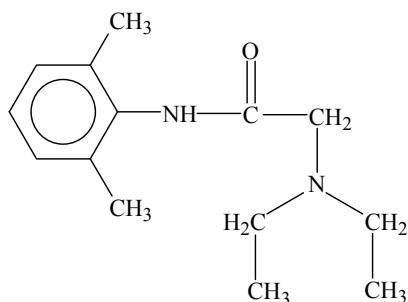
cocaine



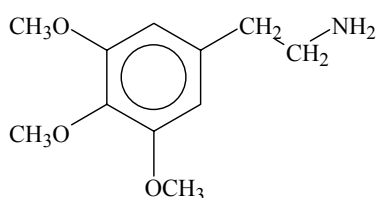
lysergic acid diethylamide (LSD)



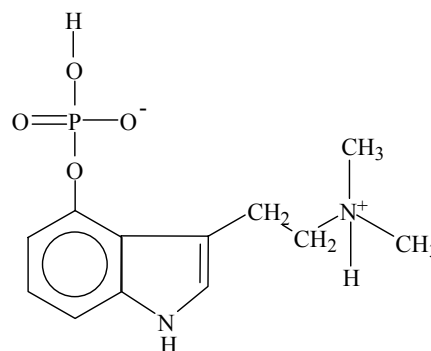
cisplatin



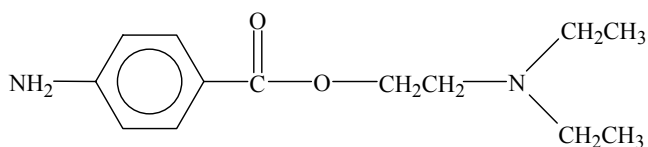
Lidocaine



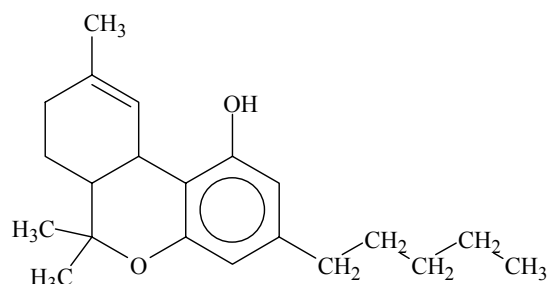
mescaline



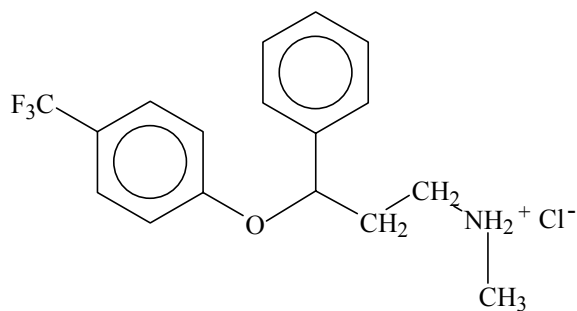
psilocybin



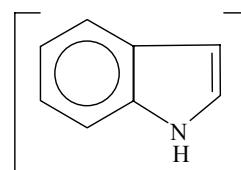
procaine



tetrahydrocannabinol (THC)

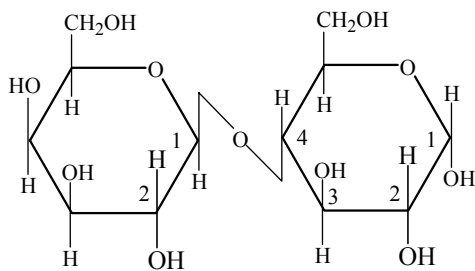


fluoxetine hydrochloride (prozac®)

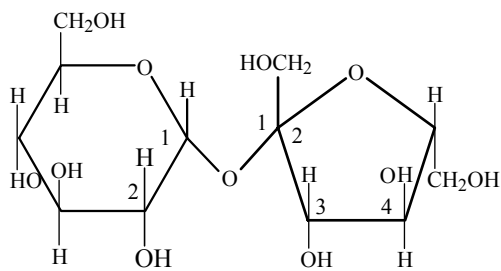


indole

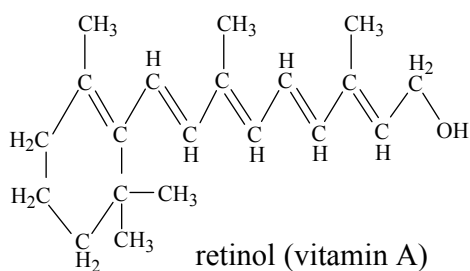
22. Structural Formulas of Some Important Biological Molecules



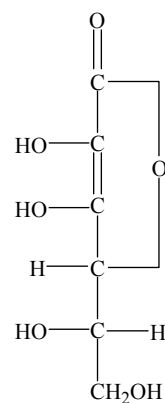
lactose



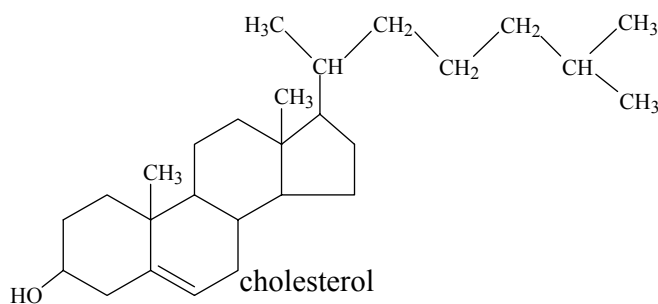
sucrose



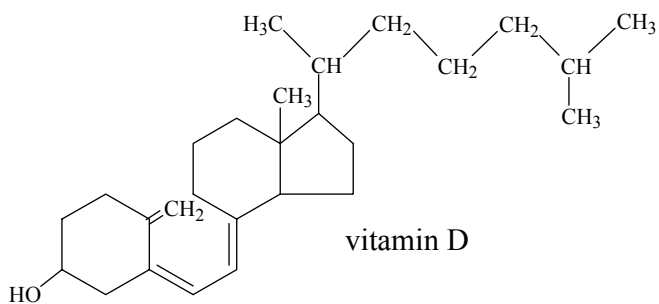
retinol (vitamin A)



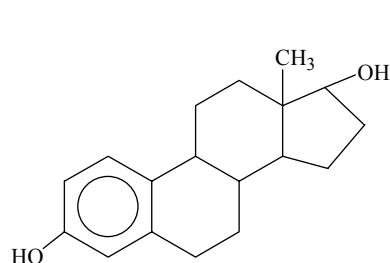
ascorbic acid (vitamin C)



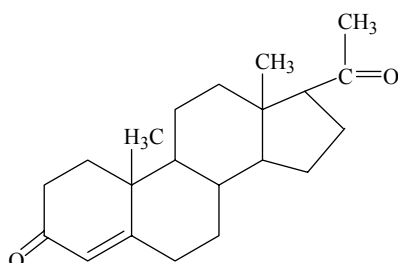
cholesterol



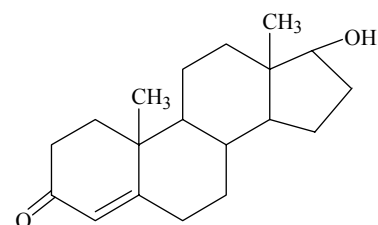
vitamin D



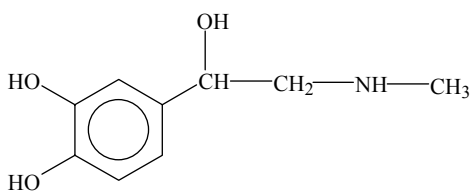
oestradiol



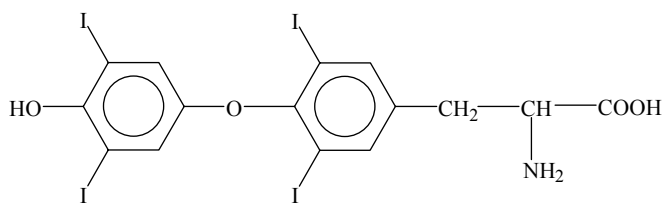
progesterone



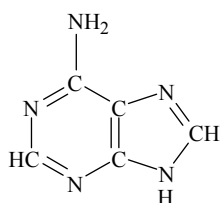
testosterone



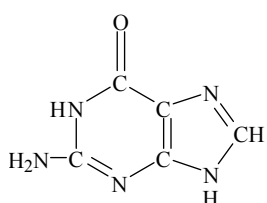
adrenaline



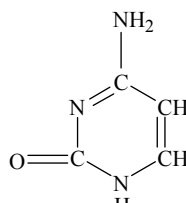
thyroxine



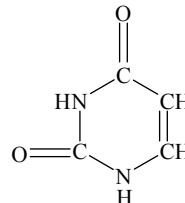
adenine



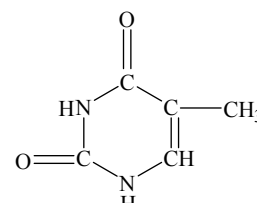
guanine



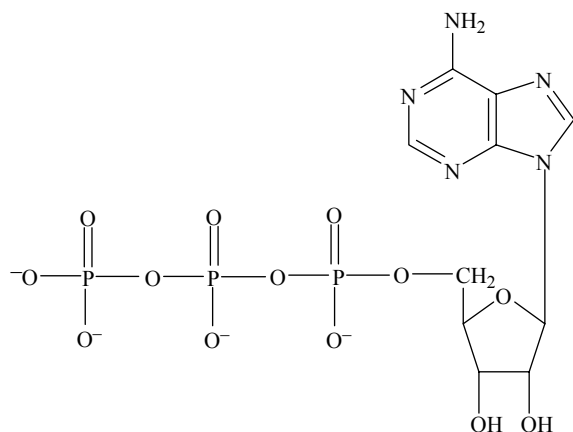
cytosine



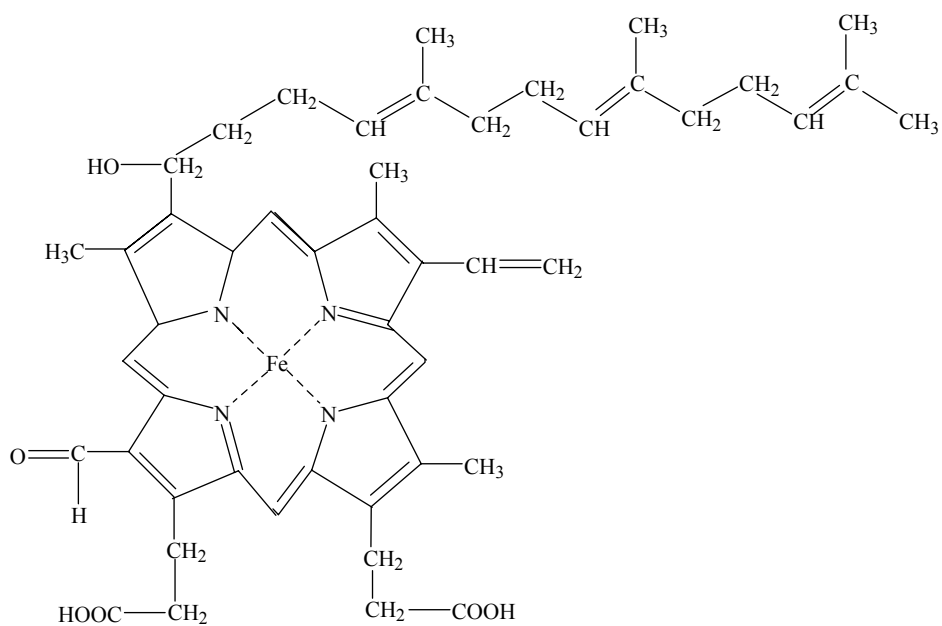
uracil



thymine



adenosine triphosphate (ATP)



The haem group from cytochrome oxidase